

The World of Chemistry: Chemical Bonding
(28 minutes)

Name _____

Period _____

READ THE QUESTIONS FIRST, BEFORE VIEWING THE VIDEO!

- _____ 1. All substances are held together by _____.
- _____ 2. The keys to bonding are the _____ and electron _____.
- _____ 3.
- _____ 4. The _____ gas elements are chemically inert.
- _____ 5. They contain filled _____ and _____ electron clouds which results in a stable electron configuration.
- _____ 6.
- _____ 7. The two types of bonding are _____ and _____.
- _____ 8.
- _____ 9. Ionic bonding results from a electrostatic _____ between ions.
- _____ 10. An example of an ionically bonded compound is _____.
- _____ 11. Salt crystals are _____ in shape.
- _____ 12. In the first demonstration ions were not free to move until the solid was _____.
- _____ 13. In the second demonstration the electrolyte conducts because the salt was dissolved in _____.
- _____ 14. Another example of an ionic salt other than sodium chloride is _____.
- _____ 15. Covalent bonding is a result of electrons being _____ between the atoms of the compound.
- _____ 16. Ionic substances are nearly all _____ (phase) where as covalent substances can be either solid, liquid or gases.
- _____ 17. In covalent substances electrons are shared by overlapping _____ clouds.
- _____ 18. When covalent bonds are formed energy is _____.
- _____ 19. The nitrogen molecule, N_2 , has _____ (number) pairs of electrons in the triple bond.
- _____ 20. Nitrogen fixers are bacteria used by _____ to convert molecular nitrogen into ammonia.